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Chapter 5 @Vedantu Young

Wonders ~~Acids Bases And~~

Salts Answers

Solution: The reaction of  
the acid + base gives a  
product of salt + water,  
which is considered as  
neutralization reaction.

Examples:  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl}$   
 $+ \text{H}_2\text{O}$

~~NCERT Solutions for class 10  
chapter 2 Acids, bases and~~

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Science Chapter 2 Acids  
Bases and Salts with Answers  
Solutions. Acids Bases and  
Salts Extra Questions Very  
Short Answer Type. Question  
1. Two solutions A and B  
have pH values of 5 and 8  
respectively. Which solution  
will be basic in nature ...

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Ans:- Salts are classified  
into four categories on the  
basis of the strength of  
acid and base. they are:  
Salt of strong acid and

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Strong base. e.g. NaCl, KCl;  
Salt of strong acid and weak  
base. e.g.  $\text{NH}_4\text{Cl}$ ,  $\text{NH}_4\text{I}$ ;  
Salt of weak acid and strong  
base. e.g.  $\text{CH}_3\text{COONa}$ ,  $\text{COO}^-$   
Na + Salt of weak acid and  
weak base. e.g.  $\text{CH}_3\text{COO}^-$   
 $\text{NH}_4^+$  + How is common salt  
found in nature?

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Bases & Oxides; 8.2 Salts &  
Chemical Analysis; 9. The  
Periodic Table ...

~~Acids, Bases & Oxides | CIE  
IGCSE Chemistry | MCQ &  
Answers~~

2. Bases react with . acids  
to produce salts and water;  
water to produce acids and  
salts; salts to produce  
acids and water; neither  
acids, salts, nor water

~~Acids Bases and Salts  
Multiple Choice Questions  
and Answers~~

The salts of strong acids  
and weak bases give acidic  
solution having pH less than

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7. Example,  $\text{NH}_4\text{Cl}$ ,  
Ammonium Chloride will have  
pH less than 7. The salts of  
weak acids and strong bases  
give basic solution having  
pH more than 7. Example,  $\text{Na}_2\text{CO}_3$ ,  
Sodium Carbonate  
will have pH more than 7.  
Question 7.

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10 Important Questions with  
...~~

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## ~~Acids And Bases Answer Key~~ ~~Teacher Worksheets~~

Answer A reaction in which  
an acid and base react with  
each other to give a salt  
and water is termed as  
neutralization reaction. For  
Example: (i)  $\text{NaOH} + \text{HCl} \rightarrow$   
 $\text{NaCl} + \text{H}_2\text{O}$  (ii)  $\text{HNO}_3 +$   
 $\text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$  15. Give  
two important uses of  
washing soda and baking  
soda. Answer Two important  
uses of washing soda are:

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Bases ...~~

Answer. (a) HCl is acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is obtained from sea water. (b) Rock salt is the common name for the mineral "halite". Its chemical formula is NaCl. It may be white or light blue or yellow depending upon impurities present in it.

~~Acids Bases and Salts  
Chapter Wise Important  
Questions ...~~

Acids Bases And Salts Note

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~~Acids Bases And Salts Note  
Taking Answers~~

(A) inorganic acid (B)  
organic acid (C) soft  
organic acid (D) strong  
inorganic acid. Answer.  
Answer: (a) 5. When milk of  
magnesia reacts with acetic

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Education produces (A) basic salts (B) acidic salts (C) neutral salt (D) complex salt. Answer. Answer: (c) 6. Which of the following phenomena will occur when a small amount of acid is added to water ...

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## ~~Salts~~

Turmeric is a natural indicator. On adding its paste to acid and base separately, which colours would be observed. (a)

Yellow in both acid and base

(b) Yellow in acid and red

in base (c) Pink in acid and

yellow in base (d) Red in

acid and blue in base.

Answer. Answer: (b) Yellow in acid and red in base

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Question 1.

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Acids, Bases and Salts In  
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Chemistry we take a look at  
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find out the difference  
between alkalis and bases,  
how acids and alkalis react  
with one another, how

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neutralisation can produce salts and the relevance of pH values.

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Acids, bases and salts An acid is a compound that is defined by its physical and chemical properties. Acids taste sour and react with metals and polyatomic ions called carbonates. A carbonate is a charged cluster of Carbon and Oxygen atoms.

~~Acids, bases and salts. 7th  
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Acid + Base ---> Salt +  
Water. Question 3: Rain  
becomes acidic due to.  
carbon dioxide. sulphur  
dioxide. nitrogen dioxide.  
all of the above. Correct

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Option is: 4. Solution :  
Rain becomes acidic because  
of carbon dioxide, sulphur  
dioxide and nitrogen dioxide.  
They dissolve in rain drops  
to form carbonic acid,  
sulphuric acid and nitric  
acid ...

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