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Answers: Moles and Stoichiometry

Practice Problems 1) How many moles of sodium atoms correspond to 1.56×10^{21} atoms of sodium? 1.56×10^{21} atoms Na \times $\frac{1 \text{ mol Na}}{2.59 \times 10^{23} \text{ atoms Na}} = 2.59 \times 10^{-3} \text{ mol Na}$ 236.022 \times 10^{23} atoms Na 2) Determine the mass in grams of each of the following: a. 1.35 mol of Fe $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$ 1 mol Fe b. 24.5 mol O

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SOLUTION: You are asked to calculate the amount (in moles) of oxygen atoms in a given sample. You are given the mass of the oxygen sample. Use the molar mass of oxygen (1 mol O = 16.00 g) to create a conversion factor that converts mass (in grams) to amount (in mol) of oxygen. 124

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Chapter 3 Stoichiometry

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